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The Influence of Mechanical Activation on Structural Evolution of Nanocrystalline SrTiO₃ Powders

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Abstract

Structural changes caused by mechanical activation of SrTiO₃ powders were investigated using a variety of methods. Average crystallite size continuously decreased with increased activation time to around 20 nm after 120 min activation, while mesopore volume and specific surface area increased accordingly. Higher activation times lead to increased agglomeration of nanoparticles to form agglomerates of around 2 μm in size, ultimately

producing a relatively stable powder, which exhibits lower microstrain than powders activated for shorter periods of time. Raman spectroscopy shows that the behavior of TO_2 and TO_4 modes is consistent with a decrease in particle size, while behavior of the nonpolar TO_3 mode is markedly different, indicating relaxation of the inversion symmetry in polycrystalline $SrTiO_3$. UV-VIS spectra show that mechanical activation has negligible effect on $SrTiO_3$, with a slight shift caused by TiO_2 contamination due to presence of air. Other than this, the mechanical activation process preserves the chemical purity of the initial powder.

Keywords: Mechanical activation; $SrTiO_3$; microstructure; ball milling; size dependence.

1. Introduction

Materials based on strontium titanate can be found in a variety of applications, like sensors [1], catalysis [2], UV detectors [3], solar cell [4], multilayer ceramic capacitors (MLCCs) [5], etc. Strontium titanate powders are also well known photocatalysts for NO_x degradation under UV light. Both stoichiometric [6] and non-stoichiometric powders [7] have been shown to exhibit considerably higher photocatalytic activity than P25 TiO_2 powder under the same illumination conditions. The nature of chemical bonding between Sr^{2+} , Ti^{4+} and O^{2-} ions in ABO_3 cubic perovskite structure leads to a unique structure which makes it a very interesting electronic material. Functional properties of $SrTiO_3$ are strongly influenced by its crystal defects, particle size, morphology, etc., which ultimately depend on the preparation methods and conditions [8-10]. State-of-the-art application requires homogeneous micrograin ceramic materials with low impurity levels, smaller particle sizes and narrow size distribution, making alternative low-cost effective methods for large-scale production of phase-pure and

nanosized SrTiO₃ particles an important topic. While numerous methods for the production of SrTiO₃ nanopowders such as sol-gel [11], hydrothermal [12], combustion [13], polymeric precursor [14], molten salt [15] and solvothermal [16] methods have been developed, SrTiO₃ is dominantly obtained by conventional solid state reaction [17]. Most of these methods require high-purity chemicals, making them too expensive for industrial scale synthesis.

Sol-gel synthesis through alkoxide hydrolysis of strontium titanate powder produced, depending on synthesis conditions, particles of 350-800 nm in size. In addition, minor fraction of the sample contained ultra-fine particles 60 nm in size and agglomerates about 2 μm in size [18]. Hydrothermal sol-gel method allows synthesis of relatively monodisperse SrTiO₃ powder with nanoparticles in 40-100 nm range, using TiCl₄ and SrCl₂ as precursors, with reaction times of 12-48 hours at 180°C, allowing limited size control and very homogeneous sample [19]. Another method to obtain ultra-fine powders involves combustion of a dehydrated form of a strontium-titanium precursor complex, resulting in single phase SrTiO₃ nanoparticles 20-40 nm in size, produced in a single-step combustion process [20]. On the other hand, in solid state reaction SrTiO₃ powders were generally produced through mechanochemical reaction of strontium carbonate or strontium oxide with titanium dioxide powders at temperatures above 1000 °C [21, 22]. These often have undesirable characteristics, such as large size, broad size distribution, hard agglomeration, phase impurities, etc. [9]. To the best of our knowledge, there have been no reports on mechanical activation of commercial strontium titanate powders to obtain nanoparticles less than 100 nm in size. There is a report of samples synthesized using a solid state reaction, followed by milling for 10-120 min, resulting in particle sizes in 100-300 nm range [23].

In SrTiO₃ thin films, it was observed that an increase in concentration of oxygen vacancies leads to an increase in lattice parameter *a*, without a change in the crystal structure

even at relatively large concentrations of oxygen vacancies [24, 25]. This increase can be correlated with strong Coulombic repulsion between strontium and titanium cations in the understoichiometric phase and high vacancies concentrations can induce significant strain in the lattice, as much as 1-2%, although no ordered arrangement of these vacancies, regardless of their concentration has ever been observed [26]. Under stoichiometric phases of strontium titanate exhibit a significant increase in electrical conductivity and semiconductor behavior compared to the insulating stoichiometric SrTiO_3 . $\text{SrTiO}_{2.5}$ phase even exhibits metallic behavior at room temperature. Thermal stability of epitaxial SrTiO_3 films shows that they are thermally stable up to around 550°C in vacuum, where desorption of SrO occurs, followed by release of SiO and additional SrO around 850°C [27].

Taking all of this into account, this work aims to investigate the influence of mechanical activation on the microstructural characteristics of SrTiO_3 powder particles and the viability of mechanical activation as a method of production of high-purity SrTiO_3 nanocrystalline powders.

2. Experimental procedure

SrTiO_3 (99% purity, mean particle size $\leq 5 \mu\text{m}$) powder was purchased from Aldrich (USA). Mechanical activation was conducted by ball-milling the samples of 6.5g in 45-cm^3 tungsten carbide jar, with balls 5 mm in diameter (powder-to-ball ratio of 1:20), in a planetary micro mill (PULVERISETTE 7 premium line, FRITSCH). Mechanical activation was conducted for different time intervals (activation time): 0, 5, 10, 30, 60, 90 and 120 minutes, in air. The samples are labeled as STO-0 to STO-120, according to the corresponding activation time.

Morphology and microstructure of non-activated and mechanically activated powders were analyzed by a Scanning Electron Microscope (SEM), Transmission Electron Microscope

(TEM) and particle size analyzer. SEM measurements were conducted using JSM-6390 LV JEOL (25 kV), where sample tablets were crushed and covered with gold (Au). TEM measurements were conducted using JEM-1400 (120kV), where samples were prepared through sonication of a 0.5% dispersion of dry powder samples in ethanol for 75 minutes deposition onto a carbon mesh. The particle size distribution (PSA) and the nature of the agglomerate were determined by a laser particle size analyzer (Mastersizer 2000, Malvern Instruments Ltd., UK). Samples were mixed with distilled water, and an aqueous suspension was obtained after 10-min ultrasonic disaggregation.

Nitrogen adsorption-desorption measurements were performed by a Sorptomatic 1990 Thermo Electron device at $-196\text{ }^{\circ}\text{C}$. The samples were degassed at room temperature for 1h under vacuum, and then heated at $5\text{ }^{\circ}\text{C}/\text{min}$ to $110\text{ }^{\circ}\text{C}$ and kept at this temperature for at least 18 h. The obtained isotherms were analyzed using Advanced Data Processing v5.1 (Thermo Electron S. p. A.) software package, where p and p_0 represented the equilibrium and saturation pressures of nitrogen at the adsorption temperature. Brunauer-Emmet-Teller (BET) method was applied for the determination of the specific surface area (S_{BET}) of the samples, using the linear part of the nitrogen adsorption isotherms [28]. Mesopore volume was calculated according to the Barrett-Joyner-Halenda method from the desorption branch of the isotherm [29].

The X-ray powder diffraction patterns were obtained using a Philips PW-150 diffractometer with Cu-K_{α} radiation, at room temperature, at angles 2θ of $10\text{-}90^{\circ}$ with a step of 0.05° and retention time of 1s per step. Rietveld analysis was performed with full refinement using GSAS II software package [30]. Obtained values of R_{wp} (weighted residual factor) varied from 6.2% to 15.6% and the Goodness of Fit indicator was $\text{GoF}\sim 1$.

Raman spectra were taken in the spectral range of 20–1200 cm^{-1} at room temperature, with a Horiba Jobin Yvon LabRam ARAMIS Raman system, using the 633 nm line of a He-Ne laser. Data was collected using a count time of 5 seconds with five averaging cycles. The samples were measured under a microscope using a 100x objective.

The diffuse reflectance UV-Vis spectra were recorded using a Nicolet Evolution 500 spectrometer with a diffuse reflectance accessory. The reflectance (R) data were converted to pseudo-absorbance $f(R)$ using the Kubelka-Munk equation. The samples were grounded, heated overnight at 110 $^{\circ}\text{C}$ and then scanned in the range of 200-800 nm using a SRS-99-010 Labsphere Spectralon white standard as reference material.

3. Results and discussion

Fig. 1. Shows the results of SEM analysis of the powders after appropriate activation times (5, 10, 30, 60, 90 and 120 min). Non-activated SrTiO_3 is a powder composed of predominantly polygonal particles 0.2–1.2 μm in size with clear grain boundaries. After mechanical activation particle size decreased, depending on activation time, and was in the range between 70 and 390 nm, with observable particle agglomeration. For longer activation times, more than 30 minutes, activation induces stronger stress, causing the formation of a defect structure in a wider layer of powder particles and energy accumulation at the surface of powder particles. This was when formation of small grains was observed, occurring through dilation of grain boundaries within individual particles [31]. As a result of these processes, surface erosion and agglomeration was observed for the samples activated for 60 minutes and more. The agglomerates can be hard and soft, while the process of their formation is called secondary agglomeration [32]. Soft agglomerates are often formed in mechanically activated

powders, and may have a negative impact on the sintering process aimed at obtaining ceramics with good performance. Continuous mechanical activation longer than 60 minutes leads to various forms of deformation, causing the strengthening or breaking of the agglomerates. Agglomerates fracture can be due to a fatigue failure mechanism and/or due to the fragmentation of fragile agglomerates. In these samples, it can be observed that longer activation time of 90 minutes leads to dominant presence of secondary agglomerates, though particles remain constantly present. Mechanical activation for 120 minutes causes larger secondary agglomerates to break up into smaller agglomerates, while forming new ones at the same time. This is consistent with continuous deformation, cold welding and fracturing to yield a powder with a refined internal structure, but coarse and stable particle size [33].

Particles size analysis (PSA) shows the same effects observed in SEM (Fig. 2): mechanical activation led to the comminution of the initial powder particles, resulting in the shift of the distribution curve to smaller particle diameters, for powders activated for 5, 10 and 30 minutes activation. Particle segregation in the initial powder sample can also be observed. Longer mechanical activation (for 60 minutes) leads to agglomeration : agglomerates are approximately 15 μm in size, while particles are 1.3 μm , resulting in bimodal size distribution observed in PSA. Further mechanical activation (for 90 minutes) leads to comminution of these agglomerates to 2 μm in size, with larger quantity in volume than that of particles, which are approximately 0.4 μm in size. More intensive fragmentation of agglomerates is observed in samples mechanically activated for 120 minutes, when the formation of new agglomerates occurs, as observed in SEM analysis (Fig. 1.). Changes in median size with activation time are shown in Table 1. Ultimately, mechanical activation led to the comminution of the initial powder particles and a wider particle size distribution for powder activated for 120 minutes.

TEM images show that the initial powder (Fig. 3a) consists of relatively large polygonal nanoparticles, hundreds of nanometers in diameter. Mechanical activation for 5 (Figure 3b) and 10 (Figure 3c) minutes causes comminution and breakage of these nanoparticles. After 30 minutes (Figure 3d) agglomeration is observed and particle size is considerably reduced (from few tens to few hundreds of nanometers). Agglomerates increase in size after 60 (Figure 3e) and 90 minutes (Figure 3f), while after 120 minutes (Figure 3g) there is, again, evidence of comminution and breakage, indicating that the agglomerates have undergone merging and sintering, only to be ground and broken again by the mechanical activation. This is all consistent with observation in SEM and PSA.

The adsorption/desorption isotherms of non-activated and mechanically activated samples were taken for a range $0 < P/P_0 < 1$ (Fig. 4. a). These isotherms correspond to type IV isotherm [34], which indicates the presence of mesopores (2-50 nm) in these samples, with a type H₃ hysteresis loop in the range $0.18 < P/P_0 < 1$. H₃ hysteresis loop is associated with the presence of plate-like particles that form aggregates at high P/P_0 . Figure 3b. shows that mechanical activation causes a significant increase in pore volume of the SrTiO₃ powder. Specific surface area (S_{BET}) values were obtained by the BET method using the linear part of the adsorption isotherms for the interval $0.05 < P/P_0 < 0.35$. Changes in the pore volume and specific surface area are shown in Fig. 4b. Non-activated powder has a specific surface of $2.05 \text{ m}^2 \text{ g}^{-1}$, and with activation, this increases up to $7.4 \text{ m}^2 \text{ g}^{-1}$ for 60 minutes of activation. After further mechanical activation, the specific surface area remained relatively constant, which can be attributed to particle agglomeration due to an increase in the particle surface energy during comminution of the initial powder particles. Mesopore volume follows similar trend, and this can be attributed to the following: during mechanical activation of up to 60 minutes, the dominant contribution to the increase in pore volume was formation of

interparticle pores, due to formation of larger particles during secondary agglomeration; further mechanical activation leads to fragmentation of agglomerates and reduction in interparticle pores, resulting in a decrease in mesopore volume. Figure 4c shows the pore size distribution of mesopores in non-activated and mechanically activated samples. The most frequent pore diameter in the non-activated sample is about 2 nm, a value close to the limit for microporous materials (<2 nm). Mechanical activation increases the average pore size and results in narrower pore size distribution.

Fig. 5. represents XRD patterns of the non-activated and activated SrTiO₃ powders. XRD analysis indicates that the mechanical activation did not effect significant change in the cubic structure of the initial powder [JCPDS Card No. 89-4934 for SrTiO₃, 84-1778 for SrCO₃, 89-4921 for TiO₂-a]. Diffraction patterns show sharp and intensive reflections for the initial powder, while increased activation times lead to an increase in peak broadening. This can be attributed to particle refinement, reduced crystallite size and generation of stress field associated with particle fragmentation and amorphization [35, 36]. After 30 minutes of mechanical activation, new phases: orthorhombic strontium carbonate (SrCO₃) and tetragonal titanium dioxide (TiO₂), were observed, which can be attributed to a reaction between the SrTiO₃ powder and atmospheric CO₂.

Mechanical activation is an intense process and leads to a significant increase in defects (vacancies, dislocations and unit cell volume). Lattice parameter a ($=b=c$), percentage distribution of the SrTiO₃, crystallite size and microstrain in mechanically activated samples were obtained using Rietveld refinement of XRD patterns (Table 1, Fig. 6).

At the beginning of the activation process, the dominant effects were attrition or erosion of the particle surface, which led to the crystallite size reduction during the first 10 minutes of mechanical activation, from 75 nm to 56 nm, with an increase in lattice microstrain

from 0.02 % to 0.04 %. Increase in lattice parameter suggests increase in the concentration of oxygen vacancies [24, 25], while relatively minor increase in microstrain suggests that these vacancies are formed primarily on the surface, which is consistent with the process of mechanical activation. Due to relatively large crystallite size, the crystal lattice is able to accommodate surface defects without significant strain. After activation for 30 minutes, crystallite size decreases to 34 nm and there is a sharp increase in microstrain to 0.14%, coupled with a decrease in lattice parameter a . This suggests that, due to smaller size, crystal structure of nanoparticles begins to strain, in order to compensate for the existence of surface defects and the increased surface free energy. Extension of the activation time to 60, 90 and 120 minutes leads to the formation of agglomerates, in order to reduce the surface free energy, and crystallite sizes are in range from 34 nm to 20 nm. The values of lattice strain vary from 0.09% to 0.26 % with maximum of 0.26 % for 90 minutes of activation for the particle size of 34nm. This suggests that mechanical activation for longer than 60 minutes caused sintering (as indicated by the increase in crystallite size after 90 minutes), where, after 90 minutes, the process is probably incomplete, resulting in higher concentration of oxygen vacancies and other defects, as indicated by the increase in lattice parameter a and microstrain. After 120 minutes, these processes are most likely complete, resulting in stabilization of crystallites through size reduction, where less stable crystals or agglomerates are split into two or more smaller but more stable crystals. All of this is consistent with SEM and TEM measurements. It is significant that the microstrain in the final sample (after 120 minute activation) is considerably lower than in any other sample activated for more than 10 minutes, suggesting that prolonged activation likely lead to stabilization of particles through repetition of particle size reduction and grain sintering, resulting in a more stable sample, in spite its relatively small crystallite size. Therefore, applied mechanical activation can

successfully lead to the formation of cubic nanocrystalline SrTiO₃ powder with particle size of around 20 nm.

Raman spectroscopy measurements were used to determine the lattice vibrational spectra of the SrTiO₃ powders. It is well known that strontium titanate has an ideal cubic perovskite structure at room temperature with $Pm\bar{3}m$ space group and vibration modes $4F_{1u}+F_{2u}$. One of the F_{1u} vibration modes is acoustic, while F_{2u} is an inactive ("silent") mode, resulting in 3 IR-active modes: $3F_{1u}$, which are Raman inactive. Therefore, no first-order Raman scattering is expected to occur in an ideal perovskite structure, instead, the Raman spectrum corresponds to the second-order scattering [37]. However, the presence of oxygen vacancies and other defects, strain effects or external electric field breaks the central symmetry of SrTiO₃, resulting in the appearance of first-order modes in bulk STO even at the 300 K [37, 38].

Raman peaks of non-activated and mechanically activated powders were recorded in the range from 20 cm⁻¹ to 1200 cm⁻¹ (Fig. 7). In accordance with literature data, two very broad Raman effects from the second-order scattering were observed from 220 cm⁻¹ to 500 cm⁻¹ and 590 cm⁻¹ to 760 cm⁻¹, respectively [37]. Both of these consist of a large number of modes, typical of a rough-grained SrTiO₃ powder, ceramics or film at 300 K. The mode observed at 1026 cm⁻¹ is also a second-order mode. Peaks observed at: 174 cm⁻¹, 540 cm⁻¹ and 790 cm⁻¹, in the non-activated powders, are assigned to following first-order modes: a doublet (LO_1+TO_2), a TO_4 mode and a (LO_4+A_{2g}) doublet, respectively [39, 40]. (E_g+B_{1g}) peak at 152 cm⁻¹, originating from R-point at the border of Brillouin zone, is barely noticeable. The origin of the modes at ~ 107 cm⁻¹ and at ~ 126 cm⁻¹ marked as X_1 and X_2 on the graph, is not entirely clear. Having in mind the analyses of other authors, X_1 and X_2 , can be attributed to

the existence of ferroelectric domains induced by defects in the crystal lattice [41], such as oxygen vacancies [42].

In the spectrum of the SrTiO₃ powder activated up to 10 min, slight changes can be observed: an increase in intensity of the (LO₁+TO₂) peak which had a characteristic "Fano" shape, decrease and widening of the mode at 250 cm⁻¹ and the amplification of the TO₄ mode and R-structural mode. The interference between the polar TO₂ mode and the wide low-energy two-phonon continuum, caused by the polarization fluctuations in polar micro- or nano-regions usually induced by defects, and results in the Fano-shape of the (LO₁+TO₂) peak [38, 43]. The intensity of TO₂ mode is proportional to the total volume of polar micro-regions. The appearance of TO₂ phonons in SrTiO₃ is usually accompanied by the appearance of remnant polarization [44]. A considerable increase in the intensity of the polar TO₄ mode also indicates an increase in the volume of the local ferroelectric domains, which coexist with the dominant paraelectric phase [39].

The intensity of the entire Raman-effect in the range from 220 cm⁻¹ to 400 cm⁻¹ generally decreases with the further increase in activation time, which especially refers to the second-order modes at higher wavenumbers. For the longest activations (90 and 120 min), this wide Raman-effect becomes even lower than the (LO₁+TO₂) peak, which is accompanied with a considerable decrease in the intensity of the second order Raman-mode in the range from 590 to 760 cm⁻¹. On the other hand, the intensity of R- mode (E_g+B_{1g}), (LO₁+TO₂) doublet and TO₄ mode continuously increases. The rise in intensity of X₁ and X₂ modes can also be noticed, especially for activation up to 60 min. Although the behavior of X₁ and X₂ modes at activation times up to 60 min is analogous to that of the TO₂ (within the doublet LO₁+TO₂) and TO₄ mode, their position does not suggest they can be assigned as first-order fundamental Raman-modes, even though they can be linked to the disorder of the SrTiO₃

lattice. Contribution of LO_4 mode is also more evident for longer activation times, while the LO_3 mode becomes noticeable only for activation over 60 min. In the second-order modes, for the activation time of 60 minutes a very weak silent doublet (LO_2+TO_3) is observed at $\sim 258\text{ cm}^{-1}$, indicating the relaxation of the inversion symmetry in polycrystalline $SrTiO_3$ [44] with increase in defect concentration and microstrain. Although the spectra of the sample activated for 120 min exhibits a slight increase of the silent peak, it still remains weak, indicating that there is no long-range structural distortion, as well as a small number of polar micro-regions and no overlap between the adjacent polar micro-regions [44].

Raman spectroscopy gave a clear evidence about the blue shift of the TO_4 mode, X_1 mode and (LO_1+TO_2) doublet. The X_1 mode shifts from 107 cm^{-1} to 109.8 cm^{-1} . The doublet (LO_1+TO_2) shifts from 174.2 cm^{-1} to 176.2 cm^{-1} for activation time up to 90 min and the analogue behavior is observed for the polar TO_4 mode, which moves from 539.5 cm^{-1} to 542.7 cm^{-1} in the spectra of powders activated for 30 - 90 min, and to 542 cm^{-1} in the sample activated for 120 min. The observed shift of polar TO_4 and TO_2 modes is primarily caused by the introduction of microstrain into the $SrTiO_3$ lattice, as well as by the reduction in crystallite size, while changes in intensity with increase of activation time are analogous to the reported change of their intensity with decrease in temperature [44]. Similar behavior of TO_4 and TO_2 modes, in both intensity and position are in accordance with the reported changes with the decrease in particle size [45]. On the other hand, the R-structural mode exhibits a shift towards lower wavenumbers (red shift) with increase in activation time.

The occurrence of the strongest $SrCO_3$ Raman line in activated powders (after 10 minutes) and increase of its intensity with prolonged activation confirms the formation of surface $SrCO_3$, also observed by XRD measurements, caused by the enhanced surface activity of activated powders and by the fact that the activation of powders was performed in air.

Optical properties of SrTiO₃ non-activated and mechanically activated samples are investigated via Uv-VIS diffuse reflectance spectra, Fig. 8. Kubelka-Munk function is applied to convert the diffuse reflectance into the equivalent band gap energy.

$$F(R_{\infty}) = (1 - R_{\infty})^2 / 2 R_{\infty} = k/s \quad (1)$$

here R_{∞} is the reflectance of the sample. Fig. 9. Shows the plot of $[F(R_{\infty}/hv)]^{1/2}$ vs. hv , in which linear segments are extended to intersect with the hv -axis to obtain the indirect band gap of the non-activated and mechanically activated samples for 120 minutes. There is a small and gradual shift in the band gap, from 3.21 to 3.19 eV, as a consequence of mechanical activation, while the experimental value of the band gap for bulk strontium titanate is 3.21-3.25 eV [46]. There are two possible sources of this shift: the increase in the amount of TiO₂ with increase in the mechanical activation time, and the increase in oxygen vacancy concentration. TiO₂ exhibits somewhat smaller band gap than SrTiO₃ [47] and its presence on the surface of the nanoparticles could lead to a gradual decrease in the overall band gap of the system. Also, ab initio calculations of vacancy defects in strontium titanate show that there is a slight decrease in the band gap with increase in oxygen vacancy concentration to 8.3% [48]. Considering that the shift towards smaller band gap in our system occurs continuously with increase in activation time, while the changes in lattice parameters indicate fluctuations in the concentration of oxygen vacancies, there is considerably stronger correlation of the band gap shift with increase in TiO₂ content than with increase in vacancy concentration.

4. Conclusion

Effect of mechanical activation on microstructural properties of SrTiO₃ powders was investigated using a wide array of methods. It was observed that prolonged mechanical

activation caused a decrease in average crystallite size from 75 nm to 20 nm, and a corresponding increase in mesopore volume and specific surface area, with maximum $S_{\text{BET}} = 7.4 \text{ m}^2/\text{g}$. Other than the appearance of SrCO_3 and TiO_2 phases caused by the presence of air during mechanical activation, no other impurities originating either from the planetary micro mill or any other source have been observed in either XRD or Raman spectra, suggesting that this method preserves the chemical purity of the initial powder. Mechanical activation causes particle fragmentation, with an increased share of large diameter particles, due to secondary agglomeration during activation. The narrowest size distribution is observed in the powders mechanically activated for 30 minutes, when aggregation of nanoparticles was first observed in SEM and TEM. Applied mechanical activation increased the intensity of all first-order Raman modes observed in the non-activated powders, and caused the appearance of new modes induced by defects and increased microstrain. The increase in intensity of $(\text{LO}_1 + \text{TO}_2)$ doublet and TO_4 mode indicated an increase in the total volume of polar micro-regions, i.e. an increase in the volume of local ferroelectric domains coexisting with the dominant paraelectric phase. Changes in the intensity of polar TO_4 and TO_2 modes with the increase in activation time is analogous to the reported changes in their intensity with temperature decrease. The observed shift in the position of polar TO modes to higher wavenumbers (blue shift) is mainly a consequence of the introduction of microstrain into the SrTiO_3 powder, which increases from 0.02 to 0.26%, as well as the consequence of a decrease in crystallite size. In contrast, markedly different behavior in terms of change in mode position was observed for the R-structural mode. In addition, mechanical activation increases the volume of ferroelectric domains in SrTiO_3 nanoparticles. UV-Vis measurements indicate that TiO_2 contamination due to activation in air causes a slight shift in the band gap towards lower values.

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Figure 1. SEM images of SrTiO₃ powder: a) non-activated; and activated for: b) 5 min, c) 10 min d) 30 min, e) 60 min, f) 90 min g) 120 min.

Figure 2. Particle size distribution of non-activated samples and samples mechanically activated for 5, 10, 30, 60, 90 and 120 min; volume cumulative frequencies shown below.

Figure 3. TEM images of SrTiO₃ powder: a) non-activated; and activated for: b) 5 min, c) 10 min d) 30 min, e) 60 min, f) 90 min g) 120 min.

Figure 4. a) N₂ adsorption/desorption isotherms for the mechanically activated samples: STO-0, STO-60, STO-90 and STO-120; b) pore size distribution of mesoporous non-activated and mechanically activated powders: STO-5, STO-10, STO-30, STO-60, STO-90 and STO-120; c) mesopore volume and specific surface area as a function of activation time.

Figure 5. XRD patterns of non-activated and activated SrTiO₃ powders.

Figure 6. Changes in crystallite size and lattice strain with activation time.

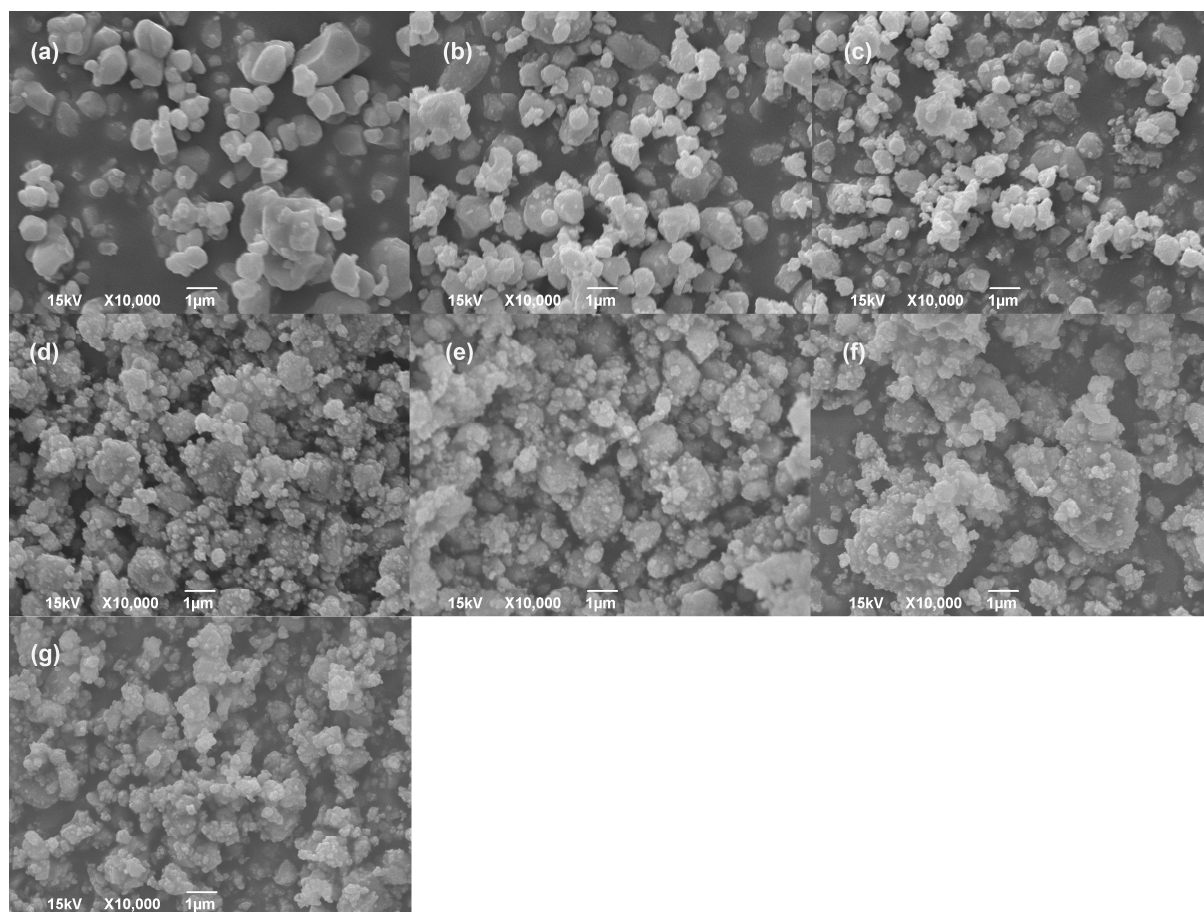
Figure 7. The influence of the mechanical activation time on the Raman-spectra of SrTiO₃ powders (assignment is described in detail in the text).

Figure 8. UV-Vis diffuse reflectance spectra for SrTiO₃ samples with different time of mechanical activation; inset: enlarged region of the onset of absorption.

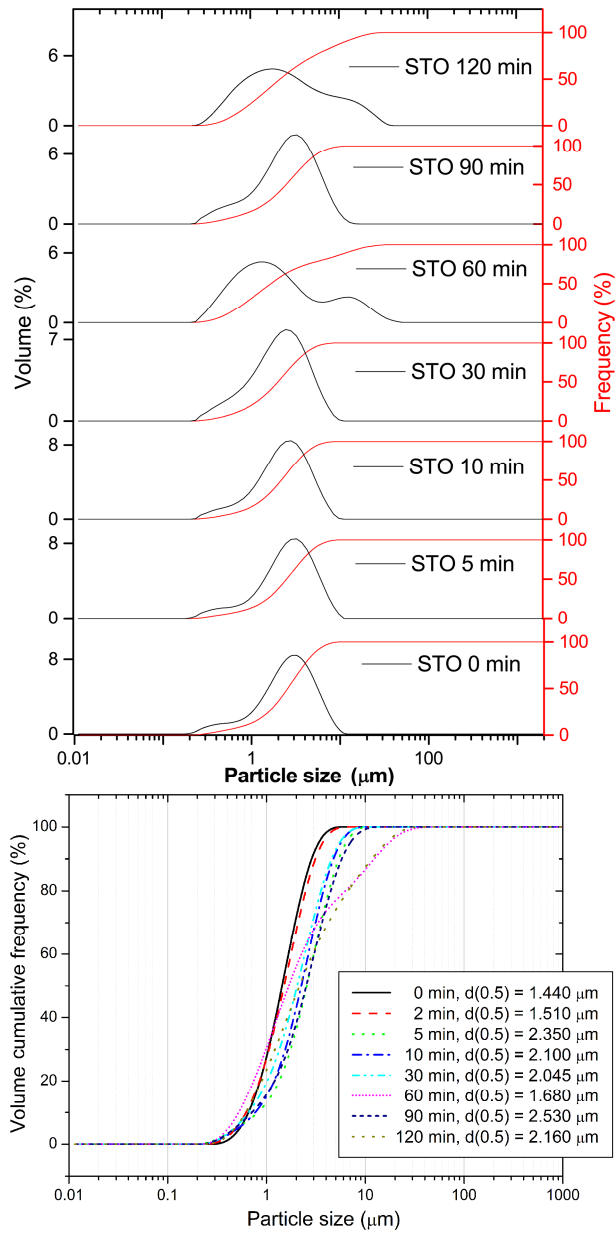
Figure 9. left: Dependence of the value of the band gap on the mechanical activation time; right: Kubelka-Munk plots and band gap energy estimation of SrTiO₃ samples: non-activated and mechanically activated in 120 minutes.

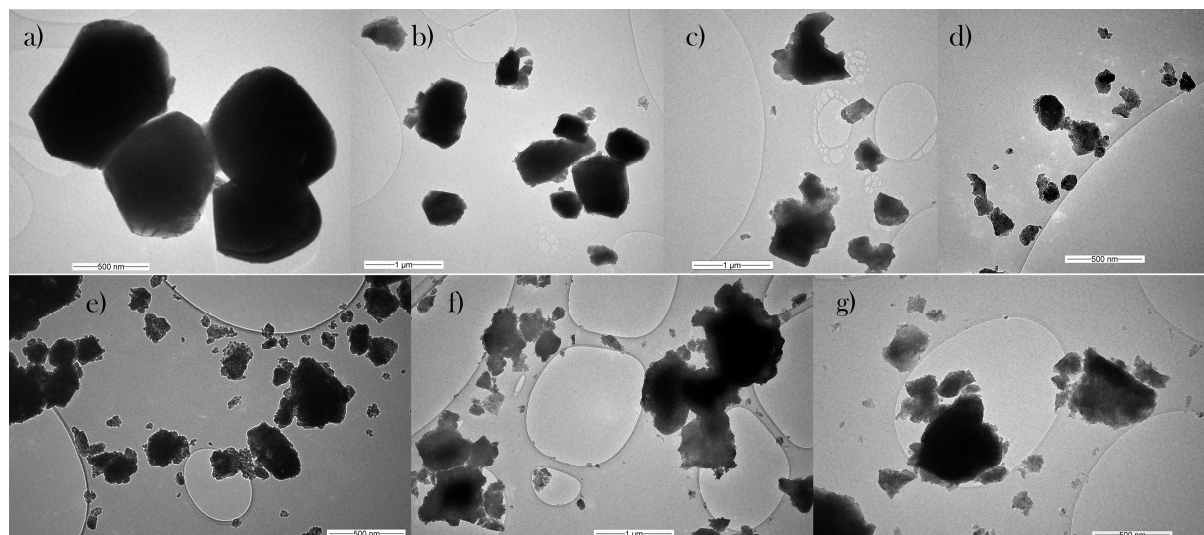
Table 1. Microstructural parameters of SrTiO₃ obtained using Rietveld, PSA and BET analysis.

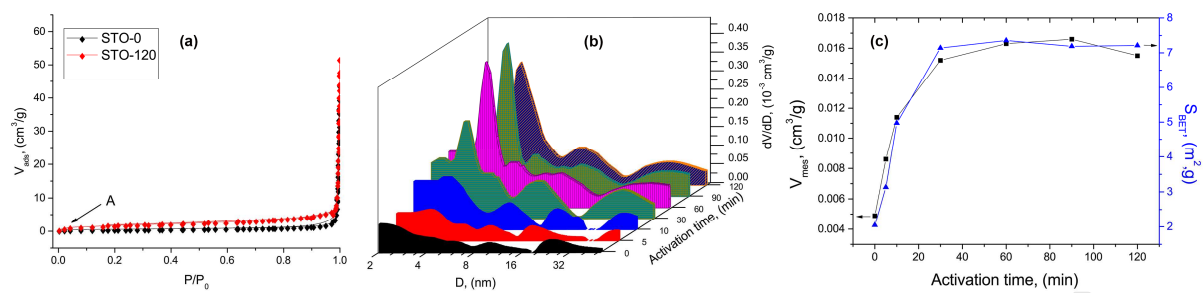
Activation time (min)	SrTiO ₃ lattice parameters (Å)	Phase content (%)		Median particle size (µm)	S _{BET} (m ² g ⁻¹)	V _{mes} (cm ³ g ⁻¹)
		SrTiO ₃ TiO ₂	SrCO ₃			
0	3.9057±0.0001	100 0	0	3.31	2.0471	0.0047
5	3.9085±0.0001	100 0	0	3.31	3.1335	0.0075
10	3.9102±0.0001	100 0	0	2.88	4.9739	0.0123
30	3.9079±0.0001	96±1 1±1	3±1	2.51	7.1419	0.0207
60	3.9066±0.0001	95±1 1±1	5±1	1.26	7.354	0.0228
90	3.9085±0.0001	92±1 2±1	6±1	3.31	7.189	0.0226
120	3.9067±0.0001	92±1 2±1	6±1	1.90	7.2104	0.0229

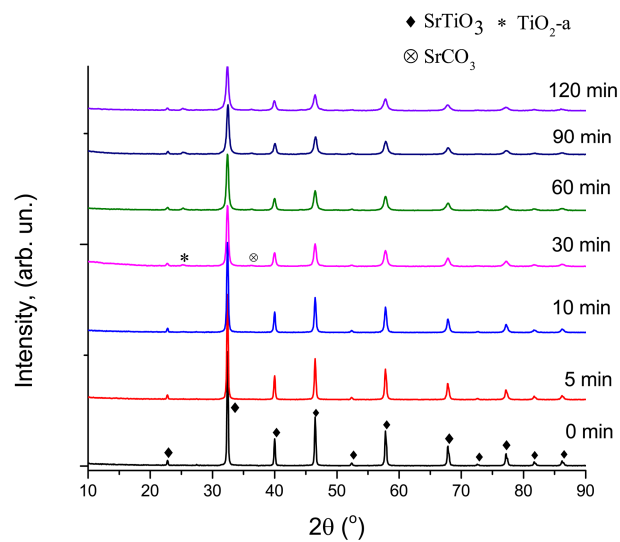


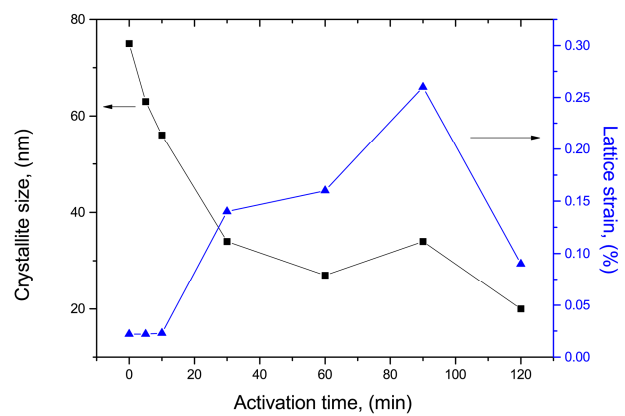
ACCEPTED MANUSCRIPT

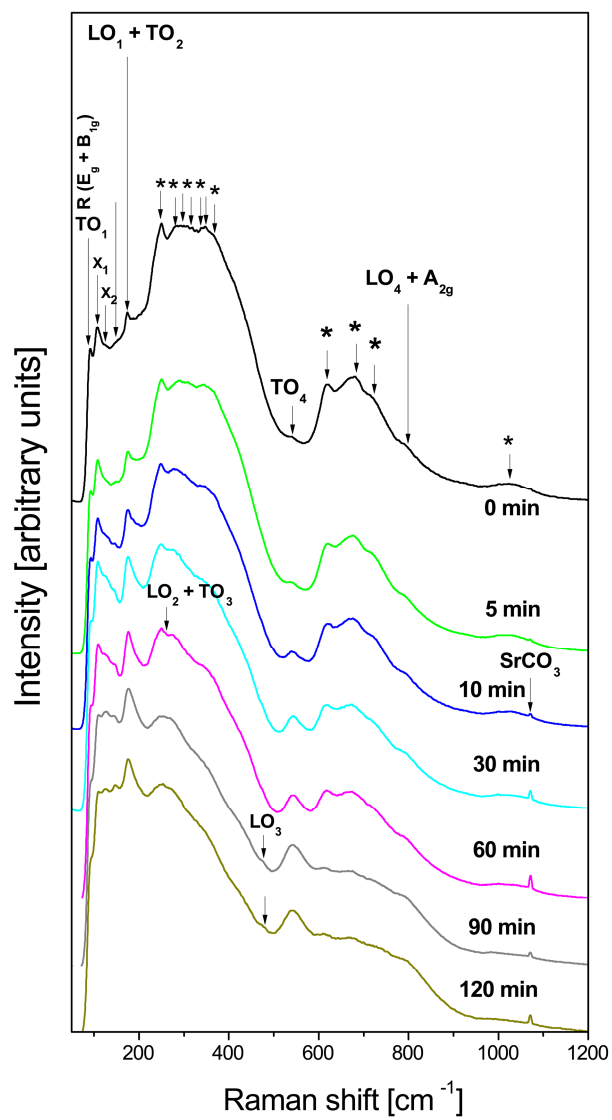


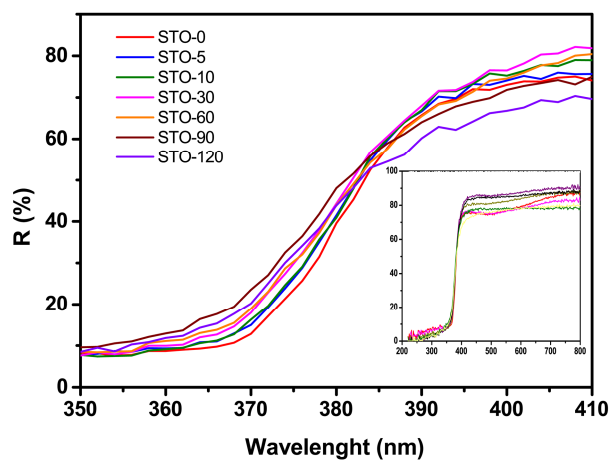


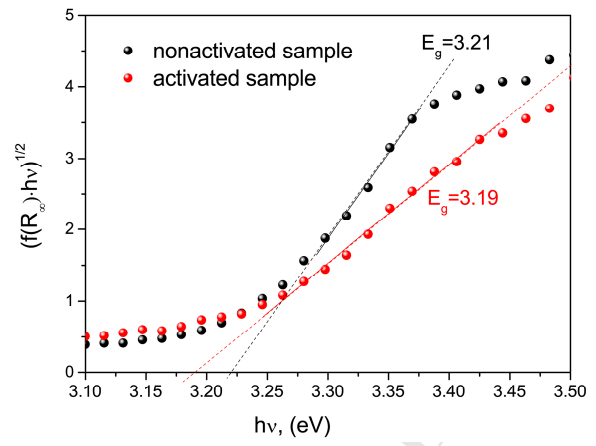
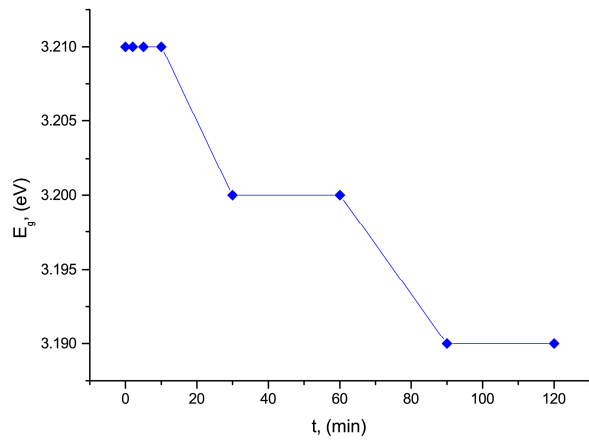












- SrTiO₃ nanocrystalline powders were prepared using mechanical activation.
- Powders mechanically activated for 30 minutes have narrowest size distribution.
- Shift in TO polar modes to higher frequencies due to introduction of micro-strains
- Observed relaxation of the inversion symmetry in SrTiO₃ due to activation
- Band gap decreases due to TiO₂ contamination, due to milling in air